



General Certificate of Education (A-level)
June 2011

Chemistry

CHEM2

(Specification 2420)

Unit 2: Chemistry In Action

Final

Mark Scheme

Mark schemes are prepared by the Principal Examiner and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation events which all examiners participate in and is the scheme which was used by them in this examination. The standardisation process ensures that the mark scheme covers the candidates' responses to questions and that every examiner understands and applies it in the same correct way. As preparation for standardisation each examiner analyses a number of candidates' scripts: alternative answers not already covered by the mark scheme are discussed and legislated for. If, after the standardisation process, examiners encounter unusual answers which have not been raised they are required to refer these to the Principal Examiner.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of candidates' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

Further copies of this Mark Scheme are available from: aqa.org.uk

Copyright © 2010 AQA and its licensors. All rights reserved.

Copyright

AQA retains the copyright on all its publications. However, registered centres for AQA are permitted to copy material from this booklet for their own internal use, with the following important exception: AQA cannot give permission to centres to photocopy any material that is acknowledged to a third party even for internal use within the centre.

Set and published by the Assessment and Qualifications Alliance.

Question	Marking Guidance	Mark	Comments
1(a)	M1 The activation energy is the <u>minimum</u> / <u>least</u> / <u>lowest energy</u> M2 (energy) <u>for a reaction</u> to occur / to go / to start OR (energy) for a <u>successful</u> / <u>effective collision</u>	2	Mark independently Ignore “heat” and ignore “enthalpy” Ignore “breaking the bonds”
1(b)	M1 Catalysts provide an alternative route OR an alternative mechanism OR alternative / different path(way) M2 Lowers the activation energy	2	Mark independently Ignore reference to “surface”
1(c)(i)	Stay(s) the same	1	
1(c)(ii)	Increases	1	Credit “increase” or “increased”
1(c)(iii)	Increases	1	Credit “increase” or “increased”
1(c)(iv)	Stay(s) the same	1	

1(d)(i)	M1 yeast or zymase M2 <u>ethanol</u>	2	Ignore “enzyme” In M2, ignore “alcohol” and ignore any formula
1(d)(ii)	M1 (Concentrated) H_3PO_4 OR (Concentrated) H_2SO_4 M2 <u>butan-2-ol</u>	2	Credit correct names Ignore “hydrogenphosphate or hydrogensulfate” Ignore “dilute” or “aq” Do not penalise absence of hyphens in name. In M2, ignore any formula

Question	Marking Guidance	Mark	Comments
2(a)	<p><i>One from</i></p> <ul style="list-style-type: none"> • Ti is not produced • TiC / <u>carbide</u> is produced OR titanium reacts with carbon • Product is brittle • Product is a poor engineering material 	1	<p>Penalise “titanium carbonate”</p> <p>Ignore “impure titanium”</p> <p>Credit “titanium is brittle”</p>
2(b)	<u>Heat (energy) change at constant pressure</u>	1	QoL
2(c)	The <u>enthalpy change</u> in a reaction is independent of the route taken (and depends only on the initial and final states)	1	Credit “heat change at constant pressure” as an alternative to “enthalpy change”

2(d)	<p>M1 The <u>enthalpy change</u> / <u>heat change at constant pressure</u> when <u>1 mol</u> of a compound / substance / product</p> <p>M2 is formed from its (constituent) <u>elements</u></p> <p>M3 with <u>all reactants and products</u> / <u>all substances in standard states</u></p> <p>OR all <u>reactants and products</u> / <u>all substances in normal states under standard conditions</u> / 100 kPa / 1 bar <u>and</u> any specified T (usually 298 K)</p>	3	<p>For M1, credit correct reference to molecule/s or atom/s</p> <p>Ignore reference to 1 atmosphere</p>
2(e)(i)	<p>Na / it is not in its <u>standard state</u> / <u>normal state under standard conditions</u></p> <p>OR</p> <p><u>Standard state</u> / <u>normal state under standard conditions</u> for Na is solid / (s)</p>	1	<p>QoL</p> <p>Ignore “sodium is a liquid or sodium is not a solid”</p>

Question	Marking Guidance	Mark	Comments
3(a)	Ca(OH)_2 OR Mg(OH)_2	1	Ignore name Could be ionic
3(b)	NaF or sodium fluoride OR NaCl or sodium chloride	1	Either formula or name can score Do not penalise the spelling “flouride” When both formula and name are written, <ul style="list-style-type: none"> penalise contradictions if the attempt at the correct formula is incorrect, ignore it and credit correct name for the mark unless contradictory if the attempt at the correct name is incorrect, ignore it and credit correct formula for the mark unless contradictory
3(c)	NaClO OR NaOCl	1	Ignore name (even when incorrect) The correct formula must be clearly identified if an equation is written
3(d)	Br₂ (ONLY)	1	Only the correct formula scores; penalise lower case “b”, penalise upper case “R”, penalise superscript Ignore name The correct formula must be clearly identified if an equation is written

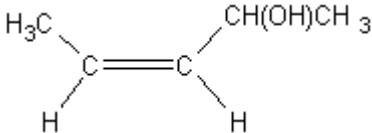
3(e)	M1 S OR S ₈ OR S ₂ M2 I ₂ (ONLY)	2	Ignore names penalise lower case “i” for iodine, penalise superscripted numbers Mark independently The correct formula must be clearly identified in each case if an equation is written
3(f)(i)	CH ₃ CH ₂ CH=CH ₂	1	Structure of but-1-ene. Ignore name Credit “sticks” for C-H bonds
3(f)(ii)	CH ₃ CH ₂ CH ₂ CH ₂ OH	1	Structure of butan-1-ol. Ignore name Credit “sticks” for C-H bonds
3(f)(iii)	CH ₃ CH ₂ CH ₃	1	Structure of propane. Ignore name Ignore calculations and molecular formula Credit “sticks” for C-H bonds Ignore the molecular ion
3(f)(iv)	CH ₃ CH ₂ Br or C ₂ H ₅ Br	1	Structure of bromoethane. Ignore name and structure of nitrile Credit “sticks” for C-H bonds

Question	Marking Guidance	Mark	Comments
4(a)(i)	$2\text{MoS}_2 + 7\text{O}_2 \longrightarrow 2\text{MoO}_3 + 4\text{SO}_2$ OR $\text{MoS}_2 + 3\frac{1}{2}\text{O}_2 \longrightarrow \text{MoO}_3 + 2\text{SO}_2$	1	Allow multiples Ignore state symbols
4(a)(ii)	<p>M1 Environmental problem</p> <p>Acid rain</p> <p>OR</p> <p>An effect either from acid rain or from an acidic gas in the atmosphere.</p> <p>M2 Use</p> <p>SO₂ could be used <u>to make / to form / to produce</u> (or wtte) H₂SO₄ / sulfuric acid</p> <p>OR</p> <p><u>To make / to form / to produce</u> (or wtte) gypsum / CaSO₄ or plaster of Paris / plaster board</p>	2	<p>Ignore references to the greenhouse effect</p> <p>Penalise reference to the ozone layer using the list principle</p>
4(a)(iii)	$\text{MoO}_3 + 3\text{H}_2 \longrightarrow \text{Mo} + 3\text{H}_2\text{O}$	1	Allow multiples Ignore state symbols

4(a)(iv)	<p>One from</p> <p>H₂ is</p> <ul style="list-style-type: none"> explosive (in)flammable <u>easily</u> ignited 	1	Ignore “burns”
4(b)(i)	<p>To allow <u>ions</u> to <u>move</u> (when molten)</p> <p>OR</p> <p><u>Ions cannot move in the solid</u></p>	1	
4(b)(ii)	$\text{Ca}^{2+} + 2\text{e}^{-} \longrightarrow \text{Ca}$	1	<p>Or multiples</p> <p>Ignore state symbols</p> <p>Ignore charge on the electron unless incorrect and accept loss of two electrons on the RHS</p>
4(b)(iii)	(High) <u>electricity</u> / <u>electrical energy</u> (cost)	1	Ignore “energy” and ignore “current”

Question	Marking Guidance	Mark	Comments
5(a)(i)	$\text{Cu} + 4\text{HNO}_3 \longrightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{NO}_2 + 2\text{H}_2\text{O}$	1	Or multiples Ignore state symbols
5(a)(ii)	M1 HNO_3 (+) 5 M2 NO_2 (+) 4	2	Ignore working out M1 Credit (V) M2 Credit (IV)
5(a)(iii)	$\text{HNO}_3 + \text{H}^+ + \text{e}^- \longrightarrow \text{NO}_2 + \text{H}_2\text{O}$ OR $\text{NO}_3^- + 2\text{H}^+ + \text{e}^- \longrightarrow \text{NO}_2 + \text{H}_2\text{O}$	1	Or multiples Ignore state symbols Ignore charge on the electron unless incorrect and accept loss of electron on the RHS
5(b)(i)	In either order M1 <u>Concentration(s) (of reactants and products) remain(s) constant / stay(s) the same / remain(s) the same / do(es) not change</u> M2 <u>Forward rate = Reverse / backward rate</u>	2	For M1 accept [] for concentration NOT “equal concentrations” and NOT “concentration(s) is/are the same” NOT “amount” Ignore “dynamic” and ignore “speed” Ignore “closed system” It is possible to score both marks under the heading of a single feature

5(b)(ii)	<p>M1 The (forward) reaction / to the right is <u>endothermic</u> or <u>takes in / absorbs heat</u> OR The reverse reaction / to the left is <u>exothermic</u> or <u>gives out / releases heat</u></p> <p>M2 depends on correct M1 and must refer to temperature/heat The <u>equilibrium shifts / moves</u> left to right to <u>oppose the increase in temperature</u></p>	2	<p>M2 depends on a correct statement for M1</p> <p>For M2, the <u>equilibrium shifts/moves</u> to <u>absorb the heat</u> OR to <u>lower the temperature</u> OR to <u>cool the reaction</u></p>
5(b)(iii)	<p>M1 refers to number of moles There are <u>fewer moles</u> (of gas) on the left OR <u>more moles</u> (of gas) on the right. OR there is <u>one mole</u> (of gas) on the left <u>and 2 moles</u> on the right.</p> <p>M2 depends on correct M1 and must refer to pressure The <u>equilibrium shifts / moves</u> right to left to <u>oppose the increase in pressure</u></p>	2	<p>M2 depends on a correct statement for M1</p> <p>For M2, the <u>equilibrium shifts/moves</u> to <u>lower the pressure</u>.</p>

Question	Marking Guidance	Mark	Comments
6(a)	<u>Pentan-2-one</u>	1	ONLY but ignore absence of hyphens
6(b)	<u>Functional group</u> (isomerism)	1	Both words needed
6(c)(i)		1	<p>Award credit provided it is obvious that the candidate is drawing the Z / <u>cis isomer</u></p> <p>The group needs to be CHOHCH₃ but do not penalise poor C–C bonds or absence of brackets around OH</p> <p>Trigonal planar structure not essential</p>
6(c)(ii)	<p>Restricted <u>rotation</u> (about the C=C)</p> <p>OR</p> <p>No (free) <u>rotation</u> (about the C=C)</p>	1	

6(d)	<table><tr><td>M1 Tollens' (reagent) <i>(Credit ammoniacal silver nitrate OR a description of making Tollens')</i> <i>(Do not credit Ag^+, AgNO_3 or $[\text{Ag}(\text{NH}_3)_2^+]$ or "the silver mirror test" on their own, but mark M2 and M3)</i></td><td>M1 Fehling's (solution) / Benedict's <i>(Penalise $\text{Cu}^{2+}(\text{aq})$ or CuSO_4 but mark M2 and M3)</i></td></tr><tr><td>M2 <u>silver mirror</u> OR <u>black solid or black precipitate</u></td><td>M2 <u>Red solid/precipitate</u> <i>(Credit <u>orange</u> or <u>brown solid</u>)</i></td></tr><tr><td>M3 (stays) colourless OR no (observed) change / no reaction</td><td>M3 (stays) blue OR no (observed) change / no reaction</td></tr></table>	M1 Tollens' (reagent) <i>(Credit ammoniacal silver nitrate OR a description of making Tollens')</i> <i>(Do not credit Ag^+, AgNO_3 or $[\text{Ag}(\text{NH}_3)_2^+]$ or "the silver mirror test" on their own, but mark M2 and M3)</i>	M1 Fehling's (solution) / Benedict's <i>(Penalise $\text{Cu}^{2+}(\text{aq})$ or CuSO_4 but mark M2 and M3)</i>	M2 <u>silver mirror</u> OR <u>black solid or black precipitate</u>	M2 <u>Red solid/precipitate</u> <i>(Credit <u>orange</u> or <u>brown solid</u>)</i>	M3 (stays) colourless OR no (observed) change / no reaction	M3 (stays) blue OR no (observed) change / no reaction	3	<p>If M1 is blank CE = 0, for the clip</p> <p>Check the partial reagents listed and if M1 has a <u>totally incorrect</u> reagent, CE = 0 for the clip</p> <p>Allow the following alternatives</p> <p>M1 (acidified) potassium dichromate(VI) (solution); mark on from incomplete formulae or incorrect oxidation state</p> <p>M2 (turns) green</p> <p>M3 (stays) orange / no (observed) change / no reaction</p> <p>OR</p> <p>M1 (acidified) potassium manganate(VII) (solution); mark on from incomplete formulae or incorrect oxidation state</p> <p>M2 (turns) colourless</p> <p>M3 (stays) purple / no (observed) change / no reaction</p> <p>In all cases for M3</p> <p>Ignore "nothing (happens)"</p> <p>Ignore "no observation"</p>
M1 Tollens' (reagent) <i>(Credit ammoniacal silver nitrate OR a description of making Tollens')</i> <i>(Do not credit Ag^+, AgNO_3 or $[\text{Ag}(\text{NH}_3)_2^+]$ or "the silver mirror test" on their own, but mark M2 and M3)</i>	M1 Fehling's (solution) / Benedict's <i>(Penalise $\text{Cu}^{2+}(\text{aq})$ or CuSO_4 but mark M2 and M3)</i>								
M2 <u>silver mirror</u> OR <u>black solid or black precipitate</u>	M2 <u>Red solid/precipitate</u> <i>(Credit <u>orange</u> or <u>brown solid</u>)</i>								
M3 (stays) colourless OR no (observed) change / no reaction	M3 (stays) blue OR no (observed) change / no reaction								

6(e)(i)	Spectrum is for Isomer 1 or named or correctly identified	1	The explanation marks in 6(e)(ii) depend on correctly identifying Isomer 1. The identification should be unambiguous but candidates should not be penalised for an imperfect or incomplete name. They may say “the alcohol” or the “alkene” or the “E isomer”
6(e)(ii)	<p>If Isomer 1 is correctly identified, award any two from</p> <ul style="list-style-type: none"> (Strong / broad) absorption / peak in the range <u>3230 to 3550</u> cm^{-1} or specified value <u>in this range</u> or <u>marked correctly</u> on spectrum and (characteristic absorption / peak for) <u>OH</u> group / <u>alcohol</u> group No absorption / peak in range <u>1680 to 1750</u> cm^{-1} or absence <u>marked correctly</u> on spectrum and (No absorption / peak for a) <u>C=O</u> group / <u>carbonyl</u> group / <u>carbon-oxygen double bond</u> Absorption / peak in the range <u>1620 to 1680</u> cm^{-1} or specified value <u>in this range</u> or <u>marked correctly</u> on spectrum and (characteristic absorption / peak for) <u>C=C</u> group / <u>alkene</u> / <u>carbon-carbon double bond</u> 	2	<p>If 6(e)(i) is incorrect or blank, CE=0</p> <p>Allow the words “dip” OR “spike” OR “trough” OR “low transmittance” as alternatives for absorption.</p> <p>Ignore reference to other absorptions e.g. C-H, C-O</p>

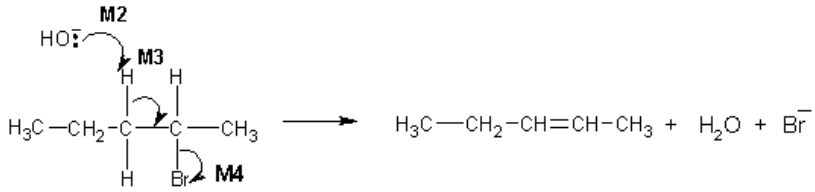
Question	Marking Guidance	Mark	Comments
7(a)(i)	$\text{CH}_4 + 3\text{F}_2 \longrightarrow \text{CHF}_3 + 3\text{HF}$	1	
7(a)(ii)	<p>M1 Initiation $\text{F}_2 \longrightarrow 2\text{F}\cdot$</p> <p>M2 First propagation $\text{F}\cdot + \text{CHF}_3 \longrightarrow \cdot\text{CF}_3 + \text{HF}$</p> <p>M3 Second propagation $\text{F}_2 + \cdot\text{CF}_3 \longrightarrow \text{CF}_4 + \text{F}\cdot$</p> <p>M4 Termination (must make C₂F₆) $2 \cdot\text{CF}_3 \longrightarrow \text{C}_2\text{F}_6 \text{ or } \text{CF}_3\text{CF}_3$</p>	4	<p>Penalise absence of dot once only.</p> <p>Radical dot on $\cdot\text{CF}_3$ can be anywhere but if the structure is drawn out, the dot must be on the carbon atom. Penalise this error once only.</p> <p>Penalise once only for a line and two dots to show a bond.</p> <p>Penalise each of “F\cdot” and lower case f, once only in this clip</p>
7(b)(i)	<p><u>Displayed formula</u> e.g.</p> <pre> Cl H Cl — C — C — F Cl F </pre>	1	<p><u>All bonds</u> must be drawn out.</p> <p>Ignore bond angles. Penalise “sticks”</p>

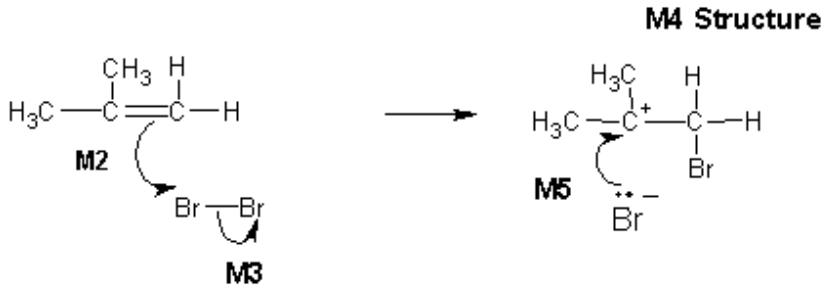
7(b)(ii)	<p>M1 <u>C—Cl</u> bond OR <u>carbon-chlorine</u> bond</p> <p>M2 <u>chlorine atom</u> OR <u>chlorine (free) radical</u></p> <p>M3 $2\text{O}_3 \longrightarrow 3\text{O}_2$</p>	3	<p>M1 NOT carbon-halogen</p> <p>Penalise incorrect spelling of chlorine <u>once only</u> in this clip</p> <p>M2 ignore formulae</p> <p>Ignore Cl_2 or $\text{Cl}\cdot$ or $\text{ClO}\cdot$ balanced on <u>both</u> sides of the equation</p> <p>Ignore other equations leading to the overall equation</p>
----------	--	---	---

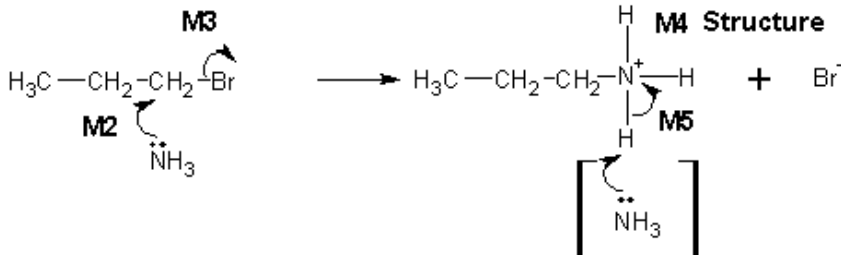
Question	Marking Guidance	Mark	Comments
8(a)(i)	<p>M1 (yellow precipitate is) silver iodide OR AgI (which may be awarded from the equation)</p> <p>M2 $\text{Ag}^+ + \text{I}^- \longrightarrow \text{AgI}$ (Also scores M1 unless contradicted)</p> <p>M3 sodium chloride OR NaCl</p>	3	<p>For M2</p> <p>Accept multiples</p> <p>Ignore state symbols</p> <p>Allow crossed out nitrate ions, but penalise if not crossed out</p>
8(a)(ii)	<p>The silver nitrate is acidified to</p> <ul style="list-style-type: none"> • react with / remove <u>ions that would interfere</u> with the test • prevent the formation of other <u>silver precipitates</u> / <u>insoluble silver compounds</u> that would interfere with the test • remove (other) <u>ions that react</u> with the silver nitrate • react with / remove carbonate / hydroxide / sulfite (ions) 	1	Ignore reference to “false positive”
8(a)(iii)	<p>M1 and M2 in either order</p> <p>M1 <u>Fluoride</u> (ion) OR <u>F⁻</u></p> <p>M2</p> <ul style="list-style-type: none"> • <u>Silver fluoride</u> / <u>AgF</u> is soluble / dissolves (in water) • <u>no precipitate</u> would form / <u>no visible / observable</u> change 	2	<p>Do not penalise the spelling “flouride”,</p> <p>Penalise “fluride” once only</p> <p>Mark M1 and M2 independently</p>

8(b)	<p>M1 $\text{Ba}^{2+} + \text{SO}_4^{2-} \longrightarrow \text{BaSO}_4$ (or the ions together)</p> <p>M2 <u>white precipitate</u> / <u>white solid</u> / <u>white suspension</u></p> <p>M3 Barium meal or (internal) X-ray or to block X-rays</p> <p>M4 <u>BaSO₄</u> / <u>barium sulfate is insoluble</u> (and therefore not toxic)</p>	4	<p>For M1, ignore state symbols</p> <p>Allow crossed out sodium ions, but penalise if not crossed out</p> <p>For M2, ignore “milky”</p> <p>If BaSO₃ OR BaS used in M1 <u>and</u> M4, penalise once only</p> <p>For M3 Ignore radio-tracing</p> <p>For M4 NOT barium ions</p> <p>NOT barium</p> <p>NOT barium meal</p> <p>NOT “It” unless <u>clearly</u> BaSO₄</p>
------	---	---	---

<p>8(c)</p>	<p>M1 <u>$2(12.00000) + 4(1.00794) = 28.03176$</u></p> <p>M2 Ethene and CO or “they” have an imprecise M_r of 28.0 / 28 OR Ethene and CO or “they” have the <u>same M_r to one d.p.</u> OR These may be shown by two clear, simple sums identifying both compounds</p> <p>M3 $\text{C}_2\text{H}_4 + 2\text{O}_2 \longrightarrow 2\text{CO} + 2\text{H}_2\text{O}$ ($\text{H}_2\text{C}=\text{CH}_2$)</p> <p>M4 <u>Displayed formula</u></p> <div style="text-align: center;"> $\begin{array}{cc} \text{H} & \text{H} \\ & \\ -\text{C} & - & \text{C}- \\ & \\ \text{H} & \text{H} \end{array}$ </div> <p>M5 Type of polymer = <u>Addition</u> (polymer)</p>	<p>5</p> <p>M1 must show working using 5 d.p. for hydrogen Penalise “similar” or “close to”, if this refers to the <u>imprecise value in M2</u>, since this does not mean “the same”</p> <p>For M3, accept $\text{CH}_2=\text{CH}_2$ OR CH_2CH_2</p> <p>For M4, <u>all bonds</u> must be drawn out including those on either side of the unit. Penalise “sticks” Ignore brackets around correct repeating unit but penalise “n”</p> <p>Penalise “additional”</p>
-------------	--	---

Question	Marking Guidance	Mark	Comments
9(a)(i)	<p>M1 Elimination</p>  <p>M2 must show an arrow from the <u>lone pair on the oxygen</u> of a negatively charged hydroxide ion <u>to a correct H atom</u></p> <p>M3 must show an arrow from a C-H bond adjacent to the C-Br bond towards the appropriate C-C bond. Only award if a reasonable attempt has been made at the attack on the H atom of the appropriate adjacent C-H</p> <p>M4 is independent provided it is from their <u>original molecule</u></p> <p>Award full marks for an E1 mechanism in which M3 is on the correct carbocation.</p> <p>N.B. These are double-headed arrows</p>	4	<p>For M1, accept “Base elimination” but <u>no other prefix</u>.</p> <p>Penalise M2 if covalent KOH</p> <p>Penalise M4 for formal charge on C of C-Br or incorrect partial charges on C-Br</p> <p>Ignore other partial charges</p> <p>Penalise once only in any part of the mechanism for a line and two dots to show a bond.</p> <p><u>Max any 2 of 3 marks for the mechanism</u> for wrong reactant (or wrong product if shown).</p> <p>Accept the correct use of “sticks” for the molecule except for the C-H being attacked</p>
9(a)(ii)	<p>Structure for pent-1-ene</p> <p><chem>CH3CH2CH2CH=CH2</chem></p>	1	<p>Penalise <chem>C3H7</chem></p> <p>Accept correct “sticks”</p>

<p>9 (b)</p>	<p>M1 <u>Electrophilic addition</u></p>  <p>M2 must show an arrow from the double bond towards the Br atom of the Br-Br molecule</p> <p>M3 must show the breaking of the Br-Br bond.</p> <p>M4 is for the structure of the tertiary carbocation with Br on the correct carbon atom.</p> <p>M5 must show an arrow from the lone pair of electrons on the negatively charged bromide ion towards the positively charged carbon atom.</p> <p>N.B. These are double-headed arrows</p>	<p>5</p> <p>For M1, both words required.</p> <p>For the mechanism</p> <p>M2 Ignore partial negative charge on the double bond.</p> <p>M3 Penalise partial charges on Br-Br bond if wrong way and penalise formal charges</p> <p>Penalise once only in any part of the mechanism for a line and two dots to show a bond</p> <p><u>Max any 3 of 4 marks for the mechanism</u> for wrong organic reactant or wrong organic product (if shown) or primary carbocation.</p> <p>If HBr is used, max 2 marks for their mechanism</p> <p>Accept the correct use of “sticks”</p>
--------------	---	--

9(c)	<p>M1 Nucleophilic substitution</p>  <p>M2 must show an arrow from the lone pair of electrons on the nitrogen atom of an ammonia molecule to the C atom.</p> <p>M3 must show the movement of a pair of electrons from the C-Br bond to the Br atom. M3 is independent provided it is from their <u>original molecule</u></p> <p>M4 is for the structure of the alkylammonium ion, which could be a condensed formula. A positive charge must be shown on/or close to, the N atom.</p> <p>M5 is for an arrow from the N-H bond to the N atom.</p> <p>Award full marks for an S_N1 mechanism in which M2 is the attack of the ammonia on the intermediate carbocation.</p> <p>N.B. These are double-headed arrows</p>	5 <p>For M1, both words required.</p> <p>Penalise M2 if NH₃ is negatively charged.</p> <p>Penalise M3 for formal charge on C or incorrect partial charges</p> <p>The second mole of ammonia is not essential for M5; therefore ignore any species here.</p> <p>Penalise once only for a line and two dots to show a bond.</p> <p>Max any 3 of 4 marks for the mechanism for wrong organic reactant (or wrong organic product if shown)</p> <p>Accept the correct use of “sticks”</p>
------	--	--

General principles applied to marking CHEM2 papers by CMI+ June 2011

It is important to note that the guidance given here is generic and specific variations may be made at individual standardising meetings in the context of particular questions and papers.

Basic principles

- **Examiners should note that throughout the mark scheme, items that are underlined are required information to gain credit.**
- **Occasionally an answer involves incorrect chemistry and the mark scheme records CE = 0, which means a chemical error has occurred and no credit is given for that section of the clip or for the whole clip.**

A. The “List principle” and the use of “ignore” in the mark scheme

If a question requires **one** answer and a candidate gives two answers, no mark is scored if one answer is correct and one answer is incorrect. There is no penalty if both answers are correct.

N.B. Certain answers are designated in the mark scheme as those which the examiner should “Ignore”. These answers are not counted as part of the list and should be ignored and will not be penalised.

B. Incorrect case for element symbol

The use of an incorrect case for the symbol of an element should be penalised **once only** within a clip. For example, penalise the use of “h” for hydrogen, “CL” for chlorine or “br” for bromine.

C. Spelling

In general

- The names of chemical compounds and functional groups **must be spelled correctly** to gain credit.
- Phonetic spelling may be acceptable for some chemical terminology.

N.B. Some terms may be required to be spelled correctly or an idea needs to be articulated with clarity, as part of the “Quality of Language” (**QoL**) marking. These will be identified in the mark scheme and marks are awarded only if the QoL criterion is satisfied.

D. Equations

In general

- Equations **must** be balanced.
- When an equation is worth two marks, one of the marks in the mark scheme will be allocated to one or more of the reactants or products. This is independent of the equation balancing.
- State symbols are generally ignored, unless specifically required in the mark scheme.

E. Reagents

The command word “Identify”, allows the candidate to choose to use **either** the name or the formula of a reagent in their answer. In some circumstances, the list principle may apply when both the name and the formula are used. Specific details will be given in mark schemes.

The guiding principle is that a reagent is a chemical which can be taken out of a bottle or container. Failure to identify complete reagents **will be penalised**, but follow-on marks (e.g. for a subsequent equation or observation) can be scored from an incorrect attempt (possibly an incomplete reagent) at the correct reagent. Specific details will be given in mark schemes.

For example, **no credit** would be given for

- the cyanide ion or CN^- when the reagent should be potassium cyanide or KCN;
- the hydroxide ion or OH^- when the reagent should be sodium hydroxide or NaOH;
- the $\text{Ag}(\text{NH}_3)_2^+$ ion when the reagent should be Tollens’ reagent (or ammoniacal silver nitrate). In this example, no credit is given for the ion, but credit could be given for a correct observation following on from the use of the ion. Specific details will be given in mark schemes.

In the event that a candidate provides, for example, **both** KCN and cyanide ion, it would be usual to ignore the reference to the cyanide ion (because this is not contradictory) and credit the KCN. Specific details will be given in mark schemes.

F. Oxidation states

In general, the sign for an oxidation state will be assumed to be positive unless specifically shown to be negative.

G. Marking calculations, such as those involving enthalpy changes

In general

- The sign for an enthalpy change will be assumed to be positive unless specifically shown to be negative.
- A correct answer alone will score **full marks** unless the necessity to show working is specifically required in the question.
- A correct numerical value with the **wrong sign** will usually score **only one mark**.

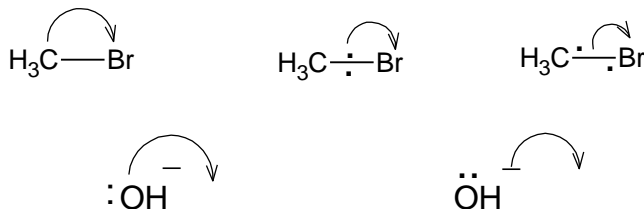
All other values **gain no credit** except

- Two marks can be awarded for correct chemistry with an arithmetic error.
- One mark can be awarded for a correct mathematical statement (or cycle) for the method.

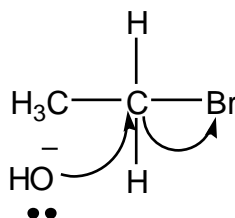
H. Organic reaction mechanisms

Curly arrows should originate either from a lone pair of electrons or from a bond.

The following representations should not gain credit **and will be penalised each time** within a clip.



For example, the following would score zero marks



When the curly arrow is showing the formation of a bond to an atom, the arrow can go directly to the relevant atom, alongside the relevant atom or **more than half-way** towards the relevant atom.

In free-radical substitution

- The absence of a radical dot should be penalised **once only** within a clip.
- The use of double-headed arrows or the incorrect use of half-headed arrows in free-radical mechanisms should be penalised **once only** within a clip

In mass spectrometry fragmentation equations, the absence of a radical dot on the molecular ion and on the free-radical fragment would be considered to be two independent errors and both would be penalised if they occurred within the same clip.

I. Organic structures

In general

- Displayed formulae must show all of the bonds and all of the atoms in the molecule, but need not show correct bond angles.
- Bonds should be drawn correctly between the relevant atoms.
For example, if candidates show the alcohol functional group as C – HO, they should be penalised **on every occasion**.
- Latitude should be given to the representation of C – C bonds in structures, given that CH₃– is considered to be interchangeable with H₃C– even though the latter would be preferred.
- Poor presentation of vertical C – CH₃ bonds or C – NH₂ bonds should **not** be penalised. For the other functional groups, such as – OH and – CN, the limit of tolerance is the half-way position between the vertical bond and the relevant atoms in the attached group.
By way of illustration, the following would apply

(a)		(b)	
	allowed		allowed
(c)		(d)	
	allowed		allowed

- In most cases, the use of “sticks” to represent C – H bonds in a structure should **not** be penalised. The exceptions will include structures in mechanisms when the C – H bond is essential (e.g. elimination reactions in haloalkanes) and when a displayed formula is required.

- Some examples are given here of **structures** for specific compounds that should **not** gain credit

CH_3COH	for	ethanal
$\text{CH}_3\text{CH}_2\text{HO}$	for	ethanol
OHCH_2CH_3	for	ethanol
$\text{C}_2\text{H}_6\text{O}$	for	ethanol
CH_2CH_2	for	ethene
$\text{CH}_2.\text{CH}_2$	for	ethene
$\text{CH}_2:\text{CH}_2$	for	ethane

N.B. Exceptions may be made in the context of balancing equations

- Each of the following **should gain credit** as alternatives to correct representations of the structures.

$\text{CH}_2 = \text{CH}_2$	for	ethene, $\text{H}_2\text{C}=\text{CH}_2$
$\text{CH}_3\text{CHOHCH}_3$	for	propan-2-ol, $\text{CH}_3\text{CH}(\text{OH})\text{CH}_3$

J. Organic names

As a general principle, non-IUPAC names or incorrect spelling or incomplete names should **not** gain credit. Some illustrations are given here.

but-2-ol	should be butan-2-ol
2-hydroxybutane	should be butan-2-ol
butane-2-ol	should be butan-2-ol
2-butanol	should be butan-2-ol
2-methpropan-2-ol	should be 2-methylpropan-2-ol
2-methylbutan-3-ol	should be 3-methylbutan-2-ol
3-methylpentan	should be 3-methylpentane
3-mythylpentane	should be 3-methylpentane
3-methypentane	should be 3-methylpentane
propanitrile	should be propanenitrile
aminethane	should be ethylamine (although aminoethane can gain credit)
2-methyl-3-bromobutane	should be 2-bromo-3-methylbutane
3-bromo-2-methylbutane	should be 2-bromo-3-methylbutane
3-methyl-2-bromobutane	should be 2-bromo-3-methylbutane
2-methylbut-3-ene	should be 3-methylbut-1-ene
difluorodichloromethane	should be dichlorodifluoromethane

K. Additional sheets and blank clips

- Markers should **mark all that is seen** and carry on marking as normal. Clips which refer to the use of additional sheets should **not** be referred to the senior team. Clips which refer to other parts of the script must be referred to the senior team.
- When considering crossed out work, **mark it** as if it were not crossed out **unless** it has been replaced by a later version; this later version then takes priority.
- Mark a blank section with a dash (—) and **not with a score of zero**.

UMS conversion calculator www.aqa.org.uk/umsconversion